

Catalytic converters are generally made from

- 1 carbon
- 2 alkaline metals
- 3 transition metals
- 4 hydrogen

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Concept:

A **catalytic converter** is a simple device that uses basic **redox reactions** to reduce the pollutants a car makes. It converts around 98% of the harmful fumes produced by a car engine into less harmful gases. *It is composed of a metal housing with a ceramic honeycomb-like interior with insulating layers. This honeycomb interior has thin wall channels that are coated with a washcoat of aluminum oxide. This coating is porous and increases the surface area, allowing more reactions to take place and containing precious metals such as **platinum**, **rhodium**, and **palladium**.* No more than 4-9 grams of these precious metals are used in a single converter.

The converter uses simple oxidation and reduction reactions to convert the unwanted fumes. Recall that oxidation is the loss of electrons and that reduction is the gaining of electrons. The precious metals mentioned earlier promote the transfer of electrons and, in turn, the conversion of toxic fumes.

Explanation:

Rhodium is one of a number of **transition metals** used in the **catalytic converters** found in diesel and petrol cars. These converters help break down polluting compounds in car engine exhaust fumes into less harmful compounds. Other metals used in catalytic converters include **palladium** and **platinum**.

** Some Hindi Fonts are not rendered properly.*

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